**Chapter 2 Review**

Read through the chapter (17 – 28), answer the questions and define key terms. This should be review!!!!! (Hint: It follows the order you will see in the chapter)

Section 2.1

Matter:

Element:

Compound:

What four elements make up about 96% of the human body?

2.2 Fill out the chart on trace elements

|  |  |  |  |
| --- | --- | --- | --- |
| Trace element | Function in the body | Where we get it? | Other information/ diseases |
| Iron |  | Diet: meat, spinach, other |  |
| Iodine |  |  |  |
| Fluoride |  |  |  |

2.3

Atom:

Proton:

Electron:

Neutron:

Nucleus:

Atomic Number:

Atomic Mass:

Isotope:

Radioactive Isotope:

2.5

Electron Shells:

2.6

Covalent Bonds:

Molecule:

Electronegativity:

Non-polar covalent:

Polar covalent:

2.7

Ion:

Ionic Bond:

2.8

Hydrogen Bond:

2.9

In the following reaction, put a square around the reactants and put a circle around the products.

6 CO2 + 6 H2O + Light --> C6H12O6 + 6 O2

2.10-2.13 Water

Adhesion:

Cohesion:

Surface tension:

Describe how water moderates temperature.

How do you know that ice is less dense than liquid water?

Solvent:

Solute:

Why is water good solvent?

2.14 Acid and Base

Acids contribute \_\_\_\_\_ (H+ or OH-) into an aqueous solution. A strong acid has a pH of \_\_\_. Bases contribute \_\_\_\_\_ (H+ or OH-) into an aqueous solution. A strong base has a pH of \_\_\_. On the pH scale, lemon juice at a pH of 2 has \_\_\_ times more H+ than an equal amount of cola at a pH of 3 and \_\_\_\_ times more H+ than tomato juice at a pH of 4.

How do buffers work?

6 CO2 + 6 H2O + Light --> C6H12O6 + 6 O2